

Course E-Syllabus

1	Course title	General chemistry lab
2	Course number	٥٥.٢١١٣
3	Credit hours	1
	Contact hours (theory, practical)	1,3
4	Prerequisites/corequisites	5502101-5502102
5	Program title	Bachelor Program in Biological sciences
6	Program code	
7	Awarding institution	Jordan university
8	School	Basic and Marine Sciences
9	Department	Biological sciences
10	Level of course	First year
11	Year of study and semester (s)	2019-2020 second semester
12	Final Qualification	Bachelor
13	Other department (s) involved in teaching the course	
14	Language of Instruction	English
15	Teaching methodology	<input type="checkbox"/> Blended <input type="checkbox"/> Online
16	Electronic platform(s)	<input type="checkbox"/> Moodle <input type="checkbox"/> Microsoft Teams <input type="checkbox"/> Skype <input type="checkbox"/> Zoom <input type="checkbox"/> Others...in Lab.....
17	Date of production/revision	

18 Course Coordinator:

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19 Other instructors:

Name:
Office number:
Phone number:
Email:

Name:
Office number:
Phone number:
Email:

20 Course Description:

As stated in the approved study plan.

In this course we will cover laboratory techniques, Molar mass of a volatile liquid, Formula of a Hydrate, Molar mass from Freezing point depression, Limiting reactant, determination of acetic acid in vinegar, : determination of an equilibrium constant, Hydrolysis, PH, and Buffers and thermochemistry and Hess's Law

21 Course aims and outcomes: A- Aims:

- To reinforce the material the student have learned the general chemistry class (5502101 and 5502102) and to give the students
- Describe the safety strategy learned in this course and how following that strategy facilitates safe behavior in everyday life.
- Describe how to properly take measurements, record data, perform calculations
- Determine fundamental physical and chemical properties of chemical compounds.

B- Intended Learning Outcomes (ILOs):

- 1- Determine the density of a solution, pure water and a solid
- 2- Determine the molar mass of unknown volatile solvent
- 3- Determine the formula of an alum and the percentage of water in anhydrate.
- 4- Determine the molar mass of an unknown solute from Freezing point depression
- 5- Determine the limiting reactant in a reaction
- 6- separate a heterogeneous mixture based on each component's physical properties

22. Topic Outline and Schedule:

Week	Lecture	Topic	Teaching Methods*/platform	Evaluation Methods**	References
1		Experiment 1 Laboratory Techniques, Techniques and Measurements			General Chemistry, by R. Change. Fifth Edition
2		Experiment 2 Molar mass of a volatile liquid			
3		Experiment 3 Formula of a hydrate		lab quiz	
4		Experiment 4 Molar mass from Freezing point depression			
5		Experiment 5 Limiting reactant			
6		Experiment 6 Physical Separation of Mixtures		lab quiz	

- Teaching methods include: Synchronous lecturing/meeting; Asynchronous lecturing/meeting

- Evaluation methods include: Homework, Quiz, Exam, pre-lab quiz...etc

23 Evaluation Methods:

Opportunities to demonstrate achievement of the ILOs are provided through the following assessment methods and requirements:

Evaluation Activity	Mark	Topic(s)	Period (Week)	Platform
reports	40	Experiment 1- Experiment 6	1-6	
Quiz	5	Experiment 3	7	
Quiz	5	Experiment 6	9	Microsoft forms

24 Course Requirements (e.g: students should have a computer, internet connection, webcam, account on a specific software/platform...etc):

Pen ,papers and lab coat

25 Course Policies:

A- Attendance policies:

B- Absences from exams and submitting assignments on time:

C- Health and safety procedures:

D- Honesty policy regarding cheating, plagiarism, misbehavior:

E- Grading policy:

F- Available university services that support achievement in the course:

26 References:

A- Required book(s), assigned reading and audio-visuals:
General Chemistry, by R. Change. Fifth Edition

B- Recommended books, materials and media:

Chemistry, by Steven S.Zumdahl, 8th edition,

27 Additional information:

Name of Course Coordinator: -----Signature: ----- Date: -----

Head of Curriculum Committee/Department: ----- Signature: -----

Head of Department: ----- Signature: -----

Head of Curriculum Committee/Faculty: ----- Signature: -----

Dean: ----- Signature: -----